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1. draw a lewis structure for each atom in the compound, making sure to place one electron on each side before pairing any electrons. 2. determine the number of valence electrons in the compound. 3. place the atom with the lowest electronegativity in the center of the atom (with the exception of Carbon). 4.

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Assessment Chapter Test A - Ed W. Clark High School

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a set of vocabulary from chapter 5 with additional terms added by the instructor Holt Chemistry NY: Chapter 5 - Ions and Ionic Compounds study guide by joel_monaco includes 20 questions covering vocabulary, terms and more.

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