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In such case  $pH = pK_a$  for acid buffer and  $pOH = pK_b$  for basic buffer. For a buffer the ratio of concentration of salt to acid or base must lie between 0.1 and 10. Thus pH range for acid buffer is from  $pK_a - 1$  to  $pK_a + 1$ . For basic buffer the pOH range is  $pK_b - 1$  to  $pK_b + 1$ .

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All right, so that's the idea of solubility and molar solubility. In part B our goal is to calculate the solubility product constant,  $K_{sp}$ , at 25 degrees Celsius for lead two chloride.  $K_{sp}$  is really just an equilibrium constant. So let's think about a solubility equilibrium. Let's think about this picture right up here.

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As, the concentration of  $H^+$  ion increases pH of water decreases, but it is still neutral. 8. A saturated solution of sodium chloride contains ionic equilibrium between dissociating sodium chloride molecules and recombining sodium chloride ions.

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So the chloride anion isn't going to react with hydronium and therefore we're not decreasing the concentration of one of our products and so our equilibrium doesn't shift and the solubility of silver chloride is unaffected by the addition of an acid. So decreasing the pH for silver chloride won't increase the solubility.

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